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RAMAN SPECTROSCOPIC STUDY OF MOLECULAR ORIENTATION IN VITREOUS --ETC(U)

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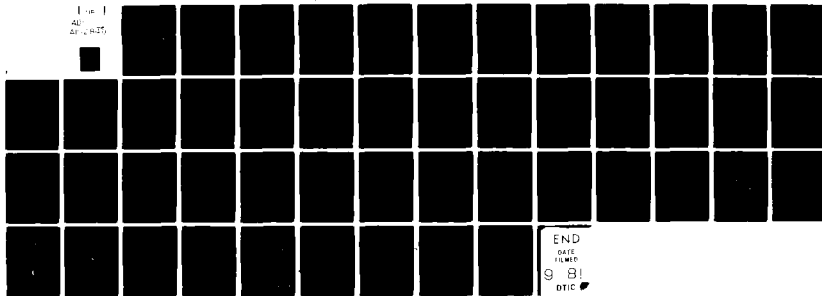
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Raman Spectroscopic Study of Molecular Orientation in Vitreous B_2O_3 Films

by

Charles F. Windisch and William M. Risen, Jr.

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Raman Spectroscopic Study of Molecular Orientation in Vitreous B_2O_3

Films

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Abstract

Preferred orientation of a molecular unit in vitreous B_2O_3 films prepared under tensile stress is demonstrated by laser Raman spectroscopy. The polarized Raman scattering is reported and found to be strongly dependent on sample orientation and polarization of the incident and scattered light. The ratios of scattering intensities in the various scattering arrangements for films are substantially different from those for bulk B_2O_3 (gl). Calculated Raman intensity ratios based upon the proposed oriented boroxol ring structure compare well with the experimental values. These results demonstrate the tendency of a molecular unit with the properties of a boroxol ring to become preferentially oriented in response to an applied stress, and provide evidence for the existence of boroxol rings in B_2O_3 (gl).

Introduction

It has been proposed widely that vitreous B_2O_3 is composed principally of interconnected boroxol (B_3O_3) rings (1-10). The postulated boroxol structure, shown in Fig. 1, is planar and has D_{3h} point group symmetry. The planarity of the boroxol ring structure distinguishes it from other possible clusters, or ways of connecting, BO_3 -triangles, and serves to allow the BO_3 -triangles, whose presence is consistent with ^{11}B -NMR findings (9), to exist in an intermediate range pseudomolecular unit.

The vibrational spectra of ^{18}O -isotopically labelled $B_2O_3(gl)$, $B_2^{16}O_3(gl)$, and mixed isotope $B_2O_3(gl)$ support the postulate that boroxol rings constitute a dominant structural feature of $B_2O_3(gl)$ (11). From that study it is also clear that the main Raman-active band of $B_2O_3(gl)$ assigned to the boroxol ring, observed at 808 cm^{-1} in $B_2^{16}O_3(gl)$, is not strongly coupled to the interconnecting network. This band is strong and relatively narrow, and in bulk samples it is highly polarized.

There have been alternative proposals for the structure which do not involve the boroxol ring, or, indeed, any intermediate range order. Thus, Elliott concluded that the observed X-ray diffraction radial distribution function can be computed at least as well by assuming a continuous random network (CRN) of planar BO_3 -units (12). Galeener, Lucovsky and Mikkelsen (10), in reporting their vibrational spectra, also have pointed out the inconsistency of conflicting interpretations of the X-ray data, and examined the feasibility of interpreting the spectra on the basis of the CRN model. Moreover, Soules and Varshneya (26) reported that their molecular dynamic calculations on B_2O_3 do not indicate boroxol ring formation if no

directional bonding is included in the potential function.

If the boroxol ring structure does exist in vitreous B_2O_3 and in the melt (or at least the material in the transformation range), it is a unit which should respond to an applied mechanical stress in a predictable manner.

When material is subjected to an applied stress, the structure will tend to respond by relaxing to a structure, or set of orientations of the components, which has a lower potential energy in the new stress field than the original structure. The kinetics of the relaxation process depends, of course, on the physical state of the material. Further, a response which involves rearrangement of intermediate range units will be faster in the liquid than the solid. Thus, if the stress is applied to a glass-forming fluid in the transformation range, the response that is effected can be quenched into the glass.

Since the spectral evidence for boroxol rings in B_2O_3 (gl) has also been observed for the melt at temperatures of $1000^{\circ}C$ and higher (13), it is reasonable to postulate that such a unit is present in B_2O_3 in the transformation region as well. If it is the type of structure suggested, it should respond to stresses applied to the melt, and the stress-induced arrangement should be quenched in the glass. A thin film of vitreous B_2O_3 , prepared by careful expansion of a bubble of the B_2O_3 melt, is formed under stress - the stress is that due to the pressure difference between the inside and outside of the bubble. This stress should induce the planar boroxol rings, if they are present, to become preferentially ordered parallel to the plane of the film.

Although it may be intuitively obvious that this would be the

expected response of a planar unit, it can be seen more clearly by considering a thin-walled spherical shell with a positive internal pressure, p_i . The shell will exhibit a stress distribution characterized by σ_r and σ_T , the stress components in the radial and tangential directions. It can be shown (14), that $\sigma_r = p_i/2$ and $\sigma_T = p_i r/2t$, where t is the thickness of the shell and r is the spherical radius. For $r \gg t$, $\sigma_T \gg \sigma_r$, so the principal stresses in the spherical film are directed tangentially, i.e. along the surface of the sphere. This situation is analogous to that of a thin film with a force field applied to the film edges directed outward but within the film plane.

Since the material, and specifically its distribution of molecular orientations, is no longer at equilibrium once the force is applied, the effect of such forces will be to rotate molecular units to positions of lower potential energy. If a boroxol ring is a flat unit hooked to neighboring units via three extraannular bonds, application of such outward-directed planar forces should cause the ring to rotate around its center and become closer to parallel to the film plane.

Determination of molecular orientation in polymeric materials has been attempted with a variety of techniques, some of which were reviewed by Wilkes (15). The application of polarized Raman scattering to this problem, particularly in determining orientation of molecular chains in fibers, has received attention as well (16-20). Miller, Exarhos and Risen (16) first employed this technique to demonstrate molecular orientation in a completely amorphous system by showing that $(PO_3)_n^{-n}$ chains are preferentially oriented with the axis of

(NaPO₃)_x (gl) fibers drawn from the melt.

The orientational properties of solids manifest themselves in the Raman spectra through the anisotropy of the polarization dependence of Raman-scattered light. Materials whose molecular units are randomly oriented exhibit Raman spectra whose band intensities can be calculated by averaging the polarizabilities over all orientations. These intensities, then, are functions of two quantities (21), the mean value $\bar{\alpha}$ and the anisotropy, γ , of the derived molecular polarizability elements α_{pq} , where $\alpha_{pq} = (\partial^2 \alpha / \partial p \partial q)_0$, and

$$\bar{\alpha} = \frac{1}{3} (\alpha_{xx} + \alpha_{yy} + \alpha_{zz}) \quad (1)$$

$$\gamma^2 = \frac{1}{2} [(\alpha_{xx} - \alpha_{yy})^2 + (\alpha_{yy} - \alpha_{zz})^2 + (\alpha_{zz} - \alpha_{xx})^2 + 6(\alpha_{xy}^2 + \alpha_{yz}^2 + \alpha_{zx}^2)] \quad (2)$$

These two quantities, $\bar{\alpha}$ and γ , are invariant under any transformation of the macroscopic coordinates of the material. That is, their values, and hence the scattering intensities, are independent of how a bulk sample is oriented in the scattering experiment.

This is not the case if complete or partial molecular ordering occurs in the sample. In such cases at least one unique axis exists. This results in Raman band intensities which depend explicitly on the geometry of the experiment, specifically on the position of the unique axis or axes in reference to the space (laboratory) coordinates of the experiment.

The prominent 808-cm⁻¹ band in the Raman spectrum of B₂O₃ (gl) has been shown to be unusual for a glassy system in its uncoupled vibrational character. Assigned to an A₁' ring-breathing mode of the

boroxol ring, it is strong, relatively sharp, and highly polarized. Thus, it provides an ideal feature for monitoring molecular orientation as a function of film orientation. Any dependence of the 808-cm^{-1} band's intensity and polarization on spatial orientation of a B_2O_3 film different from that of this band in the spectra of bulk samples will indicate that molecular (boroxol) orientation in the film is other than random. Moreover, the boroxol model requires a particular orientation-dependence of the Raman intensities, so comparison of experimental and calculated intensities of the A_1' mode of the 808-cm^{-1} band will provide evidence for or against its occurrence in B_2O_3 glass.

Thus, if the boroxol rings exist in vitreous B_2O_3 and attain preferential orientation in the film formation process, the intensities and polarization of those Raman bands assigned to boroxol ring vibrations should depend strongly on the orientation of the film in the laboratory coordinate system. Obversely, this dependence supports the existence of orientable flat structures - boroxol rings.

In this paper the polarization dependence of the Raman spectra of $\text{B}_2\text{O}_3(\text{gl})$ films with oriented boroxol rings is calculated as a function of scattering geometry, and the results of the Raman experiments on such films are reported.

Theory

In order to compute the way the intensity of the 808-cm^{-1} boroxol ring vibration in molecularly oriented films should vary with sample orientation and incident light polarization, it is necessary to consider the derived polarizability in laboratory coordinates. To do that it is useful first to define the coordinate systems.

Geometric Considerations

To specify the results of these experiments completely, three coordinate axes are defined. These are the molecular coordinates (x,y,z) , the sample (or film) coordinates $(\bar{X},\bar{Y},\bar{Z})$, and the space (laboratory) coordinates (X,Y,Z) of the Raman experiment. The molecular coordinates are those of the boroxol ring with z taken as the C_3 axis perpendicular to the plane of the ring, so x and y lie in this plane. The sample (film) coordinates are defined for the macroscopic shape with \bar{Z} taken as that perpendicular to the plane of the film. The coordinates \bar{X} and \bar{Y} are equivalent and lie in the plane of the film.

Assuming that the result of the film preparation is that the boroxol rings exist and all are parallel to the plane of the film, there is a unique relationship between the molecular and sample coordinates as shown in Fig. 2. Here z is parallel to \bar{Z} , and x and y vary arbitrarily in the $\bar{X}\bar{Y}$ plane. This relationship, with the z and \bar{Z} axes being unique, corresponds to Case I of Snyder (20), who calculated polarizabilities for a number of types of partially oriented systems.

The space (laboratory) coordinates are defined by the experimental Raman scattering geometry. As shown in Fig. 3, X is defined as the direction of propagation of the incident beam, and Z as that of the

scattered (collected) beam. The directions of polarization (electric field vector \vec{E}) of incident and scattered radiation are defined in terms of the spatial coordinates. Commonly light polarized with \vec{E} in the plane defined by the optical path is taken as V, and that normal to it as H. Thus, polarization in the XZ plane is denoted V and that in direction Y as H.

This notation is sufficient for describing experiments on the amorphous bulk glass. For films, where orientational effects are seen, the sample orientation must be specified as well. An infinite number of film orientations is possible. To specify any one of them in terms of the space (laboratory) coordinates, it is convenient to use the Euler angles and coordinate transformation, which depends on the angles θ , ϕ , and ψ . Specification of two of these, θ and ϕ , is sufficient to define completely the film's (or molecule's) orientation with respect to the space coordinates, while the third, ψ , specifies the molecule's angular variation within the film plane and is arbitrary. The coordinates and transformation matrix given by Macaulay (22) were used in these calculations, but the calculations also were done in the two other Euler systems described in references (23) and (24).

Since there are so many possible film orientations, the experiments and calculations were carried out using particular sets of orientations, or "tilts", of which three are shown in Figs. 4a, 4b and 4c. In each of these z is confined to a particular plane defined by the space coordinates as follows. By definition: Tilt 1 is that in which z is confined to the XZ plane ($\phi = 0^\circ$), and θ varies from 0° to 90° ; Tilt 2 is that in which z is confined to the XY plane ($\theta = 90^\circ$),

and ϕ varies from 0° to 90° ; and, Tilt 3 is that in which z is confined to the VZ plane ($\phi = 90^\circ$), and ψ varies from 0° to 90° . Thus, a complete specification of the following experiments on a film is given by defining the polarization directions in terms of H and V , and indicating the type and amount of film tilt in terms of ψ and ϕ .

Derivation of Raman Scattering Intensities

The problem of calculating Raman intensities for the various possible scattering geometries reduces to a determination of the various α_{pq} , the components of the derived polarizability tensor in the laboratory coordinates, in terms of the derived molecular polarizability elements α_{pq} . The intensity of Raman scattered light polarized in a particular direction is directly proportional to the square of the relevant α_{pq} .

For a bulk material with randomly oriented scattering centers, the polarizability components can be obtained by averaging the molecular polarizability components over all orientations of the molecule. The intensity results can then be expressed in terms of the $\bar{\alpha}$ and γ , given in Eqns 1 and 2. Further, it is known that for that case:

$$\overline{(\alpha_{xx}^2)} = \overline{(\alpha_{yy}^2)} = \overline{(\alpha_{zz}^2)} = \frac{1}{45} [45(\bar{\alpha})^2 + 4\gamma^2] \quad (3)$$

and

$$\overline{(\alpha_{xy}^2)} = \overline{(\alpha_{yz}^2)} = \overline{(\alpha_{zx}^2)} = \frac{1}{15} \gamma^2 \quad (4)$$

For an A_1' mode of a boroxol ring, of D_{3h} symmetry, α_{pq} is given by (25):

$$\underline{\alpha}_{pq} = \begin{pmatrix} a & 0 & 0 \\ 0 & a & 0 \\ 0 & 0 & b \end{pmatrix} \quad (5)$$

where $a = \alpha_{xx} = \alpha_{yy}$, $b = \alpha_{zz}$, and α_{xy} , α_{xz} , and α_{yz} are zero. Thus, $\bar{\gamma}$ and γ^2 take the forms $\bar{\gamma} = (2a + b)/3$ and $\gamma^2 = (a - b)^2$. In the limit of $a \gg b$, the Raman intensities relative to that for the HH mode were calculated using Eqns 3 and 4 and are given in Table I along with the experimental results for bulk B_2O_3 (gl).

To arrive at intensities for the films which can be compared to the experimental results, appropriate account of the spatial relationships must be taken. The tensor $\underline{\alpha}_{pq}$ is related to $\underline{\alpha}_{pQ}$ by

$$\underline{\alpha}_{pQ} = T \underline{\alpha}_{pq} T' \quad (6)$$

where T is the transformation matrix (22) relating the laboratory coordinates P, Q , to the molecular coordinates, p, q ; and T' is its transpose. Using $\underline{\alpha}_{pq}$ given in Eqn 5 with T yields for Eqn 6:

$$\underline{\alpha}_{pQ} = \begin{pmatrix} \alpha_{XX} & \alpha_{XY} & \alpha_{XZ} \\ \alpha_{YX} & \alpha_{YY} & \alpha_{YZ} \\ \alpha_{ZX} & \alpha_{ZY} & \alpha_{ZZ} \end{pmatrix} \quad (7)$$

the square 3 of whose elements are given in Appendix I.

The terms α_{YX} , α_{ZX} , α_{YY} , α_{ZY} give rise to the HV, VV, HH, and VH intensities respectively, and are the only components of interest here. Calculated squares of these components for the various film orientations are given in Appendix II, and in Tables II, III and IV the relevant values are given along with the experimental results for the intensities normalized to that of the HH component. All

calculated values assume that the angle of incidence of the light with the film is the same as that with well oriented boroxol rings. This assumption is discussed below.

At most only two of these intensities are predicted to be non-zero for any orientation. The calculated non-zero ratios for the various tilts are, for $a \gg b$; VV/HH for Tilt 1, HV/HH for Tilt 2, and VH/HH for Tilt 3, and are presented in Fig. 5 as a function of the sample tilt angle (θ for Tilts 1 and 3, ϕ for Tilt 2).

It is useful to note several of the assumptions that have been made in relating α_{pq} to α_{PQ} . One is that the boroxol rings are aligned with the film, so that z is parallel to \bar{z} . If this is true of all boroxol rings, the transformation holds for all of them, and, in the absence of other effects, the experimental results would be exactly as predicted by Eqn 7. Similarly if there is preferential but not complete alignment, the experimental results should follow the trends in intensity predicted by Eqn 7.

The other assumption is that the angle of incidence of the light (Raman source) with the film is the same as that with an exactly aligned boroxol ring. Since the refractive index of bulk $B_2O_3(g1)$ is about 1.48, this holds rigorously only for rings at the surface. Refraction has an effect on some of the measurements at incidence different from 0 or 90° , and on some in which the scattered light is refracted. These effects involve both the identity of the relevant elements in α_{PQ} and on the optical angles and are discussed later.

Experimental

Glass Preparation

The B_2O_3 glasses were made by first heating H_3BO_3 (reagent grade) in a Pt crucible at $1000^\circ C$ in an electric furnace for one hour. The melt was then quenched to form a glass according to the following procedures.

Bulk samples of the B_2O_3 glass were made by pouring the melt into a cylindrical stainless steel mold under a stream of dry N_2 at room temperature. Vitreous B_2O_3 films of 10-20 μm in thickness were made by first allowing the melt to cool enough to become somewhat viscous. Then a hollow pyrex tube was dipped into the melt to coat one end with molten B_2O_3 . A carefully controlled burst of N_2 was then passed through the tube so as to form a bubble slowly. While the bubble was expanding but nearly full-blown, it was pressed lightly against a stainless steel block so that a flat section of film was obtained. It was possible to obtain reproducible B_2O_3 bubbles with walls nearly as thin as permitted by the surface tension of the melt just above T_g . For some experiments films as thick as 50-100 μm were prepared.

This effort was spent making thin films to assure that the maximum degree of in-plane stress obtainable by surface-tensile-limited film formation was introduced. This was done to obtain the highest practical state of alignment of putative boroxol rings in $B_2O_3(gl)$.

Sample Handling and Raman Spectra

Because of the hygroscopic nature of $B_2O_3(gl)$, care was taken to exclude water while handling and measuring the samples. After making some series of measurements in a dry N_2 -flushed sample holder

and some within the spectrometer while the sample itself was flushed with N_2 , the following procedure was found to be sufficient to avoid hydrolysis. The sample was surrounded with N_2 during preparation and transportation to the spectrometer. The films were mounted and studied while under a stream of dry N_2 . Each set of spectra was taken within 15 minutes of sample preparation. After each set of spectra was measured, the $800\text{--}950\text{ cm}^{-1}$ region was measured to detect any scattering at 880--cm^{-1} due to boric acid formation at the surface, and rejected if any intensity at that frequency was found.

After each sample was mounted, the $700\text{--}900\text{ cm}^{-1}$ region of the Raman spectrum was taken at the four possible combinations of incident and scattering polarization, i.e. HH, HV, VH, VV. These were achieved by appropriate adjustment of the polarization rotator and analyzer. Immediately after taking the spectra in these modes, each was repeated in turn to confirm reproducibility, and the 880--cm^{-1} region was scanned to confirm the absence of Raman-detectable hydrolysis.

To obtain a spectrum of the bulk glass, the cylindrical sample was cleaved in half and positioned as shown in Fig. 6a with the laser beam impinging on the top surface about 1 mm from the cloven edge, and radiation scattered at 90° was collected from this face.

Raman spectra of the B_2O_3 (gl) films were obtained by mounting them at the appropriate orientations and collecting the light scattered at 90° to the incident beam. Due to the thinness of the films, there were different sample mounting requirements and problems for each orientation. The spectra at Tilt 1 ($\phi = 0^\circ$) were obtained using

the sample mount shown in Fig. 6b. The vitreous film was mounted over a hole drilled through a blackened metal plate. The use of the plate allowed for more accurate control of θ . Spectra were taken for $\theta = 0, 30, 45, 60$, and 90° . The spectra in Tilts 2 and 3 were obtained by tilting the film in the beam to the desired alignment. Since the mounting arrangement used in Tilt 1 measurements was not applicable to these orientations, these angles could be measured only with lower accuracy. At least three determinations were made for each angle for these cases, at the angles of $\phi = 0, 45$, and 90° for Tilt 2 ($\theta = 90^\circ$), and $\theta = 0, 45$, and 90° for Tilt 3 ($\phi = 90^\circ$), but the uncertainty in the average angular values could be as high as 5° .

As appropriate, samples were partially masked to eliminate extraneous scattering from the edges of the film, as shown in Fig. 3. In some measurements of Tilt 2, thicker films (50-100 μm) were used, since for $\theta = 90^\circ$ and $0 < \phi < 90^\circ$ it is difficult to get the source into the edge of a 10-20 μm film without reflecting some of the light from the face. With 50 μm films the focused laser beam passes through the edge of the film.

Raman spectra of the vitreous B_2O_3 samples were taken on a Jarrell-Ash 25-300 laser Raman spectrometer using the 514.5 nm line for excitation. The spectral resolution was ca. 5 cm^{-1} and the laser power employed was ca. 750 mW. The optical arrangement in the vicinity of the sample is shown in Fig. 3. Scattered light was optically scrambled after polarization analysis and before passing the entrance slit of the monochromator. Depolarization ratios of the standard CCl_4 were adequately reproduced with this spectrometer.

Results

The observed intensities of the 808-cm^{-1} Raman band of bulk vitreous B_2O_3 for the various polarization combinations are given relative to the intensity of the HH component in Table I. The observed intensities of this band for film samples for Tilts 1, 2 and 3, again relative to the strongest band observed in each film orientation, the HH polarization component, are given in Tables II, III, and IV, respectively.

The results obtained on the bulk samples are in good agreement with those calculated for complete disorder. The values for the HV, VH and VV components relative to HH are calculated to be 0.13, in the limit $a \gg b$ (i.e. with b (α_{zz}) set to zero), and they are observed to be 0.15 ± 0.03 , 0.16 ± 0.03 , and 0.11 ± 0.03 . These values, averaging 0.14 ± 0.03 , are a bit higher than the other reported value of 0.07 (13), which presumably also has an experimental error of about ± 0.03 . Together, these values confirm that the mode is highly polarized, and that if its \bar{a} and γ^2 have the form given above for a boroxol ring, b is in the range $0 \leq b \leq 0.15a$, or $a \gg b$.

The results for the polarization ratios for the $\text{B}_2\text{O}_3(\text{gl})$ films demonstrate two main points. First, the units which vibrate at 808-cm^{-1} are not randomly arrayed; they are ordered to a significant degree. Second, the ordering of these units is approximately as calculated for boroxol rings; that is, flat molecular units parallel to the surface of the film. Although these conclusions will be qualified in detail, by noting that preferential ordering does not require that all boroxol rings be exactly parallel to the surface, as the calculations have assumed, noting some optical effects not

considered so far, and comparing the detailed data with the calculated results, the data are essentially those predicted for oriented boroxol rings.

The comparisons of calculated and experimental results that lead to the main conclusions are these. The film data are very different from those for bulk $B_2O_3(g1)$. For the bulk $B_2O_3(g1)$, all components other than HH have relative intensities of only 0.14 ± 0.03 . However, as shown in Tables II, III, and IV, for the films one component other than HH becomes quite strong as the angle is varied. Moreover, the one which does become strong, VV in Tilt 1, HV in Tilt 2, and VH in Tilt 3, is the one calculated to gain intensity upon changing film orientation in the prescribed manner. In addition, with two exceptions, all of those relative intensities calculated to be zero in this idealization are 0.15 ± 0.06 . These values may be taken to be effectively zero for the purpose of comparing calculated and observed ratios. Their difference from zero results partly from the fact that not all of the units are perfectly oriented, even though the degree of orientation is high, and from some optical effects discussed below.

The two relative intensities which should also be small, but are not, are: first, those for which the HH component derives exclusively from the polarizability element b , $\phi = 90^\circ$ in Tilt 2 and $\phi = 90^\circ$ in Tilt 3, and second the VH/HH ratio at $\phi = 45^\circ$ in Tilt 2. For the former ones, where HH is given by b^2 , it is important to note that since $a \gg b$, the HH component is very weak and all of

the other components are even weaker. Taking $b \approx 0.05a$ as a reasonable estimate from the bulk data, the HH band for these orientations is expected to be only 2.5×10^{-3} as intense as it is at the other extreme angle in each relevant Tilt. This means that the ratios at issue are, relatively, 0/0.0025, and experimental noise is a significant factor since each of them is the result of dividing two very weak bands. As discussed below, the observed values can be accounted for completely in a more straightforward way. If there is an alignment error of about 3° , the expected values for the ratios are not zero, but about 0.8. Alignment errors of $1-4^\circ$ are sufficient to account for all of these observations and are reasonable estimates of the experimental uncertainty. More relevant than the errors that show up as a result of the HH elements becoming small in those cases is the very fact that they do become very small, since these HH components are calculated to arise only from the element b^2 , which is small.

While inspection of Tables II, III, and IV shows that the types of effects predicted by the model are observed, this may be clearer from the plots of the calculated and observed data in Fig. 7. There it is seen that the VV/HH ratios go through a maximum as θ is varied, while the VH/HH and HV/HH ratios do not vary.

DISCUSSION

The pre mise of this work was the following. If boroxol rings exist and have their postulated properties, then they should orient preferentially in a film formed under tensile stress. And, if they do exist with these properties and orient, the polarization dependence of a band in the Raman spectrum due to the ring-unit should have a predictable film-orientation dependence. For the ideal case in which all boroxol rings are oriented parallel to the film's surface, the predicted dependence, detailed above and in the accompanying tables and appendices, is that in each "tilt" two of the intensity ratios should be non-zero at some varied angle, while the other ratios should be zero at all angles. The two non-zero ratios are HH/HH and VV/HH in Tilt 1, HH/HH and HV/HH in Tilt 2, and HH/HH and VH/HH in Tilt 3. Experimental observations of this sort clearly would be consistent with the presence of boroxol rings. What such findings would prove, though, is that a unit with a vibrational mode at the frequency of observation (808-cm^{-1}) exists and is orientable by the application of a tensile stress, that it has at least one unique symmetry axis along which the polarizability element is different from that along the other two principal axes, and that the scattering of greatest intensity results from polarizability along the other two principal axes, which are preferentially parallel to the surface of the film.

The main experimental finding in this study is that the predicted pattern of intensity ratios is observed. This is clearest in the data for Tilts 1 and 3. However, the complete analysis of the results and the experiments themselves requires some

rather involved considerations, which take into account optical effects, the reality of less-than-ideal orientation, and experimental error.

There are several sources of experimental error, including noise and uncertainties in orientation. As mentioned earlier, noise may be important in those cases ($\psi = 90^\circ$ in Tilt 2, $\theta = 90^\circ$ in Tilt 3) where the ratio of two extremely weak bands is calculated. Uncertainties in film orientation result partly from the difficulty in precisely aligning and handling ca 20 μ -thick films of hygroscopic material, and partly from the lack of perfect flatness of the films. While considerable care was taken and the spectra were reproduced a number of times, the angles are known imprecisely. This has several ramifications, but the most important is that the zero-valued intensity ratios can take non-zero values - up to 0.06 in cases where HH is strong, and much higher where the HH intensity is ideally proportional to b^2 . For example in Tilt 1 if $\theta = 90^\circ$ and $\phi = 10^\circ$ instead of $\psi = 90^\circ$ and $\phi = 0^\circ$, the ratio HV/HH becomes 0.03 instead of zero and is somewhat closer to the observed value of 0.13 ± 0.05 . The net effect of several small angular errors plus some scattering from unaligned units explains the fact that many of the ratios calculated to be zero are, in fact, observed to be on the order of 0.1. In those special cases noted above ($\phi = 90^\circ$ in Tilt 2, and $\theta = 90^\circ$ in Tilt 3) where ideally zero valued ratios are quite large, angular uncertainty is particularly important. In them, changing the angles by only $1-4^\circ$ takes the zero valued ratios to the 0.56-0.98 observed, so it is not unreasonable to ignore their

difference from zero in comparing the observed data to those computed for the ideal case.

There are several optical effects of potential significance in these experiments. At incident angles other than normal to the entrance surface, refraction occurs. This changes the angle of incidence between the light and sub-surface units and, in the process, its direction of incident light polarization. These effects are discussed in Appendix III, in which an approximation to the effect of changing the incident angle is treated. The effect of changing the direction of light polarization is to change the form of the polarizability tensor because the \vec{E} vector projection on the molecular coordinate system is changed. However, derivation of the new tensor shows that the new terms are nearly offsetting in their effect on the intensity ratios. Another effect is that certain scattered light rays in certain orientations attempt to exit at the critical angle, θ_c , relative to the surface. As they cannot, they are "light piped". While this can affect absolute intensities, it does not have a significant impact on any of the intensity ratios. Finally, polarization-selective reflection, especially near Brewster's angle, can be important in this sort of experiment. There is one case where it may be significant here; VV/HH at $\theta = 30^\circ$ in Tilt 1, in which the HH intensity may be made lower while the VV is unaffected. This may explain why VV/HH is higher for $\theta = 30^\circ$ than for $\theta = 60^\circ$ in Tilt 1. Overall, the optical effects considered here do not invalidate the conclusion that the calculated trends are observed.

Finally, the degree of molecular orientation in the film samples

may be considered. In light of the sources of error and the complexity of the analysis and phenomena, it is not feasible to compute the degree of order with certainty. The best estimate can be made by using the measurements at Tilt 3, $\theta = 45^\circ$, VH/HH, and Tilt 2, $\theta = 45^\circ$, HV/HH, since both components of each ratio are reasonably strong bands. For each ratio, I_{ij}/I_{HH} , both I_{ij} and I_{HH} are sums of contributions from the N_1 oriented and the N_2 random population oscillators. The forms of the contributions to the total intensity in each orientation from ordered units and from a random population of oscillators are known. From them and the ratios a quite high degree of orientation is found for those units responsible for the scattering. Given this and the fact that the refractive effects do not appear to be overriding, it is reasonable to suppose that many of these centers are near the surface.

Conclusion

The orientation dependence of the polarization dependence of the Raman spectral band at 808-cm^{-1} of thin $\text{B}_2\text{O}_3(\text{gl})$ films, formed under tensile stress, is consistent with the existence of oriented boroxol rings. Since this band also appears in the spectra of B_2O_3 melts and bulk glasses, these data support the conclusion that boroxol rings exist in $\text{B}_2\text{O}_3(\text{gl})$.

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References

1. J. Goubeau and H. Keller, Z. Anorg. allg. chem. 272, 303 (1953).
2. J. Krogh-Moe, Ark, Kemi 12, 475 (1958).
3. L.A. Kristiansen and J. Krogh-Moe, Physics Chem. Glasses 9(3), 96 (1968).
4. J. Krogh-Moe, J. Non-Cryst. Solids 1, 269 (1969).
5. R.L. Mozzi and B.E. Warren, J. Appl. Crystallogr. 3, 251 (1970).
6. W.L. Konijnendijk and J. M. Stevels, J. Non-Cryst. Solids 18, 307 (1975).
7. T.W. Bril, Philips, Res. Repts. Suppl. No. 2 (1976).
8. J.P. Bronswijk and E. Strijks, J. Non-Cryst. Solids 24, 145 (1977).
9. G.E. Jellison, Jr., L.W. Panek, P.J. Bray, and G. B. Rouse, Jr., J. Chem. Phys. 66, 802 (1977).
10. F.L. Galeener, G. Lucovsky and J.C. Mikkelsen, Jr., Physical Review B 22, 3983 (1980).
11. C.F. Windisch and W.M. Risen, Jr., to be published.
12. S.R. Elliot, Philos. Mag. B37, 435 (1978).
13. G.E. Walrafen, S.R. Samanta and P.N. Krishnan, J. Chem. Phys. 72, 113 (1980).
14. N.P. Suh, A.P.L. Turner, Elements of the Mechanical Behavior of Solids, McGraw-Hill Book Co., New York, 1975, p 74.
15. G.L. Wilkes, Adv. Polymer Sci. 8, 91 (1971).
16. P.J. Miller, G.J. Exarhos and W.M. Risen, Jr., J. Chem. Phys. 59, 2796 (1973).
17. P.W. Higgs, Proc. R. Soc. A 220, 472 (1953).
18. S.W. Cornell and J.L. Koenig, J. Appl. Phys. 39, 4883 (1968).
19. D. I. Bower, J. Polym. Sci B 10, 2135 (1972).

20. R.G. Snyder, J. Mol. Spec. 37, 353 (1971).
21. E.B. Wilson, Jr., J.C. Decius and Paul C. Cross, Molecular Vibrations, McGraw-Hill Book Company, Inc., New York, 1955.
22. W. H. Macaulay, Solid Geometry, Cambridge University Press, London, 1930, pp 60-62.
23. G. Arfken, Mathematical Methods for Physicists, Academic Press, New York, 1970, pp 178-180.
24. P. Frost, Solid Geometry, MacMillan & Co., London, 1875, p 100.
25. R. Loudon, Adv. Phys. 8, 423 (1964).
26. T. F. Soules and A. K. Varshneya, J. Am. Ceram. Soc. 64, 145 (1981).

TABLE I

Calculated and Observed Raman Intensities of the 808 cm^{-1} Band of Bulk $\text{B}_2\text{O}_3(\text{gl})$ for the Various Scattering Configurations Relative to the Intensity of the Band for the HH Configuration

	HH	HV	VH	VV
Calc.	1.0	0.13	0.13	0.13
Obs.	1.0	0.15 ± 0.03	0.16 ± 0.03	0.11 ± 0.03

TABLE II

Calculated and Observed Raman Intensity Ratios for B_2O_3 (gl) Films
in Tilt 1 ($\theta = 0^\circ$)

	HH	HV	VH	VV
element ⁽¹⁾	a^2	0	0	0
90° calc ratio ⁽²⁾	1.0	0	0	0
obs. ratio	1.0	0.13±.05	0.13±.05	0.16±.07
element	a^2	0	0	$3(a-b)^2/16$
60° calc ratio	1.0	0	0	$3(a-b)^2/16a^2$
obs. ratio	1.0	0.13±.05	0.15±.06	.33±.06 ⁽³⁾
element	a^2	0	0	$(a-b)^2/4$
45° calc ratio	1.0	0	0	$(a-b)^2/4a^2$
obs. ratio	1.0	0.15±.03	0.16±.03	0.56±.06 ⁽³⁾
element	a^2	0	0	$3(a-b)^2/16$
30° calc ratio	1.0	0	0	$3(a-b)^2/16a^2$
obs. ratio	1.0	0.14±.05	0.16±.05	.44±.06 ⁽³⁾
element	a^2	0	0	0
0° calc ratio	1.0	0	0	0
obs. ratio	1.0	0.11±.03	0.13±.03	0.2±.07

(1) Element means the relevant element in the α_{PQ}^2 tensor (Appendix I)

(2) All calculated values are based on the assumption that the angle of incidence of the light with the film is the same as that with completely aligned boroxol rings (see text). The ratios are with respect to the HH intensity.

(3) Although measured values fell in a narrower range than specified in these cases, these limits were assigned after consideration of the observed ranges in all cases.

TABLE III

Calculated and Observed Raman Intensity Ratios for B_2O_3 (gl) Films
in Tilt 2 ($\theta = 90^\circ$)

	HH	HV	VH	VV
element ⁽¹⁾	a^2	0	0	0
0° calc ratio ⁽²⁾	1.0	0	0	0
obs. ratio	1.0	0.13±.05	0.13±.05	0.16±.07
element	$(3a+b)^2/16$	$3(a-b)^2/16$	0	0
30° calc ratio	1.0	$3(a-b)^2/(3a+b)^2$	0	0
element	$(a+b)^2/4$	$(a-b)^2/4$	0	0
45° calc ratio	1.0	$(a-b)^2/(a+b)^2$	0	0
obs. ratio	1.0	0.81±0.14	0.61±.14	0.2±0.1
element	$(a+3b)^2/16$	$3(a-b)^2/16$	0	0
60° calc ratio	1.0	$2(a-b)^2/(a+3b)^2$	0	0
element	b^2	0	0	0
90° calc ratio	1.0	0	0	0
obs. ratio	1.0	0.56±*(3)	0.96±*(3)	0.68±*(3)

(1) See footnote (1), Table II

(2) See footnote (2), Table II

(3) See text

TABLE IV

Calculated and Observed Raman Intensity Ratios for B_2O_3 (gl) Film
in Tilt 3 ($\phi = 90^\circ$)

ϕ		HH	HV	VH	VV
0°	element ⁽¹⁾	a^2	0	0	0
	calc ratio ⁽²⁾	1.0	0	0	0
	obs. ratio	1.0	$0.11 \pm .03$	$0.13 \pm .03$	$0.2 \pm .08$
30°	element	$(3a+b)^2/16$	0	$3(a-b)^2/16$	0
	calc ratio	1.0	0	$3(a-b)^2/(3a+b)^2$	0
	obs. ratio	1.0	$0.15 \pm .08$	0.79 ± 0.10	$0.18 \pm .05$
45°	element	$(a+b)^2/4$	0	$(a-b)^2/4$	0
	calc ratio	1.0	0	$(a-b)^2/(a+b)^2$	0
	obs. ratio	1.0	$0.15 \pm .08$	0.79 ± 0.10	$0.18 \pm .05$
60°	element	$(a+3b)^2/16$	0	$3(a-b)^2/16$	0
	calc ratio	1.0	0	$3(a-b)^2/(a+3b)^2$	0
	obs. ratio	1.0	$0.56 \pm^*(3)$	$0.96 \pm^*(3)$	$0.68 \pm^*(3)$
90°	element	b^2	0	0	0
	calc ratio	1.0	0	0	0
	obs. ratio	1.0	$0.56 \pm^*(3)$	$0.96 \pm^*(3)$	$0.68 \pm^*(3)$

(1) See footnote (1), Table II

(2) See footnote (2), Table II

(3) See text.

Appendix I

Squares of the Elements of the Polarizability Tensor α_{PQ} in Terms
of the Molecular Coordinates

The polarizability tensor in laboratory coordinates is α_{PQ} , and the squares of its elements, which are proportional to the experimental Raman intensities, are elements of the tensor α_{PQ}^2 , where:

$$\alpha_{PQ}^2 = \begin{pmatrix} i & j & k \\ 1 & m & n \\ o & p & q \end{pmatrix}$$

and,

$$i = [a(\sin^2\phi + \cos^2\phi \cos^2\theta) + b \cos^2\phi \sin^2\phi]^2$$

$$m = [a(\cos^2\phi + \sin^2\phi \cos^2\theta) + b \sin^2\phi \sin^2\theta]^2$$

$$q = (a \sin^2\theta + b \cos^2\theta)^2$$

$$j = l = (a-b)^2 \sin^2\phi \cos^2\phi \sin^2\theta$$

$$k = o = (a-b)^2 \cos^2\phi \sin^2\theta \cos^2\theta$$

$$n = p = (a-b)^2 \sin^2\phi \sin^2\theta \cos^2\theta$$

Appendix II

The Polarizability Tensors for the Three Experimental Film Sample Orientations

The experimental Raman intensities are proportional to the squares of the elements of the tensor $\underline{\alpha}_{PQ}$. As given in Appendix I, these squares are elements of the tensor $\underline{\alpha}_{PQ}^2$, where

$$\underline{\alpha}_{PQ}^2 = \begin{pmatrix} i & j & k \\ l & m & n \\ o & p & q \end{pmatrix}$$

For the three experimental arrangements, or Tilts, studied in this work, the elements of $\underline{\alpha}_{PQ}^2$ are:

Tilt 1 ($\psi = 0^\circ$)

$$i = (a \cos^2 \psi + b \sin^2 \psi)^2$$

$$m = a^2$$

$$q = (a \sin^2 \psi + b \cos^2 \psi)^2$$

$$k = o = (a-b)^2 \sin^2 \psi \cos^2 \psi$$

$$j = l = n = p = 0$$

Tilt 2 ($\psi = 90^\circ$)

$$i = (a \sin^2 \psi + b \cos^2 \psi)^2$$

$$m = (a \cos^2 \psi + b \sin^2 \psi)^2$$

$$q = b^2$$

$$j = l = (a-b)^2 \sin^2 \psi \cos^2 \psi$$

$$k = o = n = p = 0$$

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Tilt 3 ($\phi = 90^\circ$)

$$i = a^2$$

$$m = (a \cos^2 \theta + b \sin^2 \theta)^2$$

$$q = (a \sin^2 \theta + b \cos^2 \theta)^2$$

$$n = p = (a-b)^2 \sin^2 \theta \cos^2 \theta$$

$$j = l = o = k = 0$$

APPENDIX III

It is possible to describe the optical effects in these experiments more completely by including internal reflections, polarization-selective reflection, and refraction. Although each can contribute, and certain of the "light-piping" effects may be significant in determining absolute intensities in certain orientations, the most important effect to consider is refraction. Refraction of the incident ray makes the angle of incidence of the light relative to aligned scattering centers different from that relative to the entrance surface. This alters the functional dependence of the experimental scattered intensities on the various polarizability components.

From an analysis of this functional dependence, it is clear that a reasonable approximation to the overall refraction effect is to consider only the change in incident angle with respect to the scattering centers. That is because the effects of changing the functional forms of the polarizability tensor elements nearly offset one another. A simple calculation can be done within this approximation if it is assumed that the experimental scattering geometry, the angle between the refracted incident and scattered light, remains at 90° , and that refraction does not materially affect the collected fraction of scattered light. This is equivalent to the case of a rotated film with no refraction.

The "new" tilt angle, that incident on a scattering center within the sample, can be calculated using Snell's law and the bulk refractive index of B_2O_3 glass, $n = 1.48$. The effect of these changes on the calculated relative intensities ($a \gg b$) is

shown below for Tilt 1. Here, calc' refers to values calculated by including refraction-caused changes in the tilt angle, and calc refers to the relative intensities at nominal experimental angles.

Tilt 1	Nominal θ	HH/HH	HV/HH	VH/HH	VV/HH
calc	90	1.0	0	0	0
calc'		1.0	0	0	0
calc	60	1.0	0	0	0.19
calc'		1.0	0	0	0.10
calc	45	1.0	0	0	0.25
calc'		1.0	0	0	0.18
calc	30	1.0	0	0	0.19
calc'		1.0	0	0	0.23
calc	0	1.0	0	0	0
calc'		1.0	0	0	0

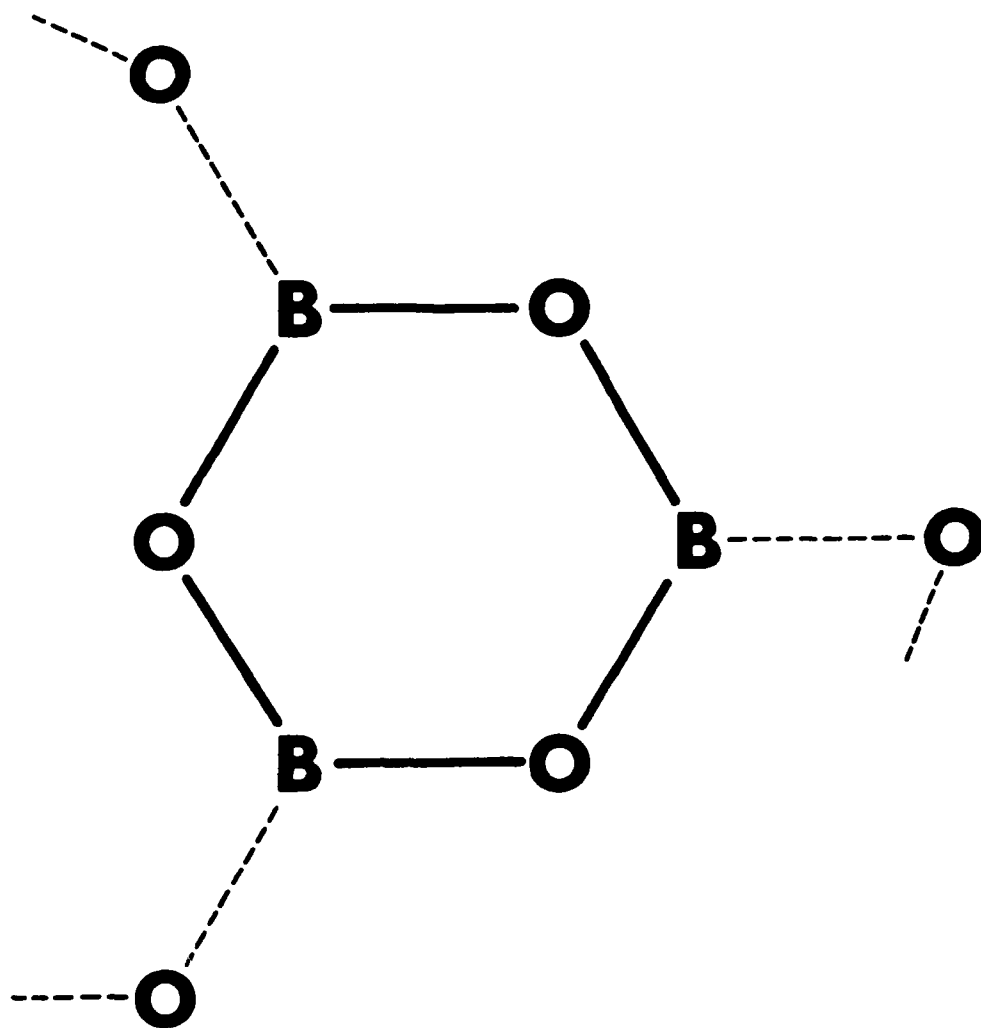
As can be seen, the refraction-effects are predicted to be significant, but they alter only the relative intensity values of the non-zero components shown (VV/HH). The zero-valued components remain zero. Thus, conclusions based on the observation of the trends (e.g. the dependence on VV/HH only in Tilt 1) are not affected by this aspect of refraction. Moreover, the predicted effect of refraction was not observed in these experiments; rather the data followed the more idealized calculations. This may mean that most of the scattering centers involved in these measurements were at or near the surface.

Figure Captions

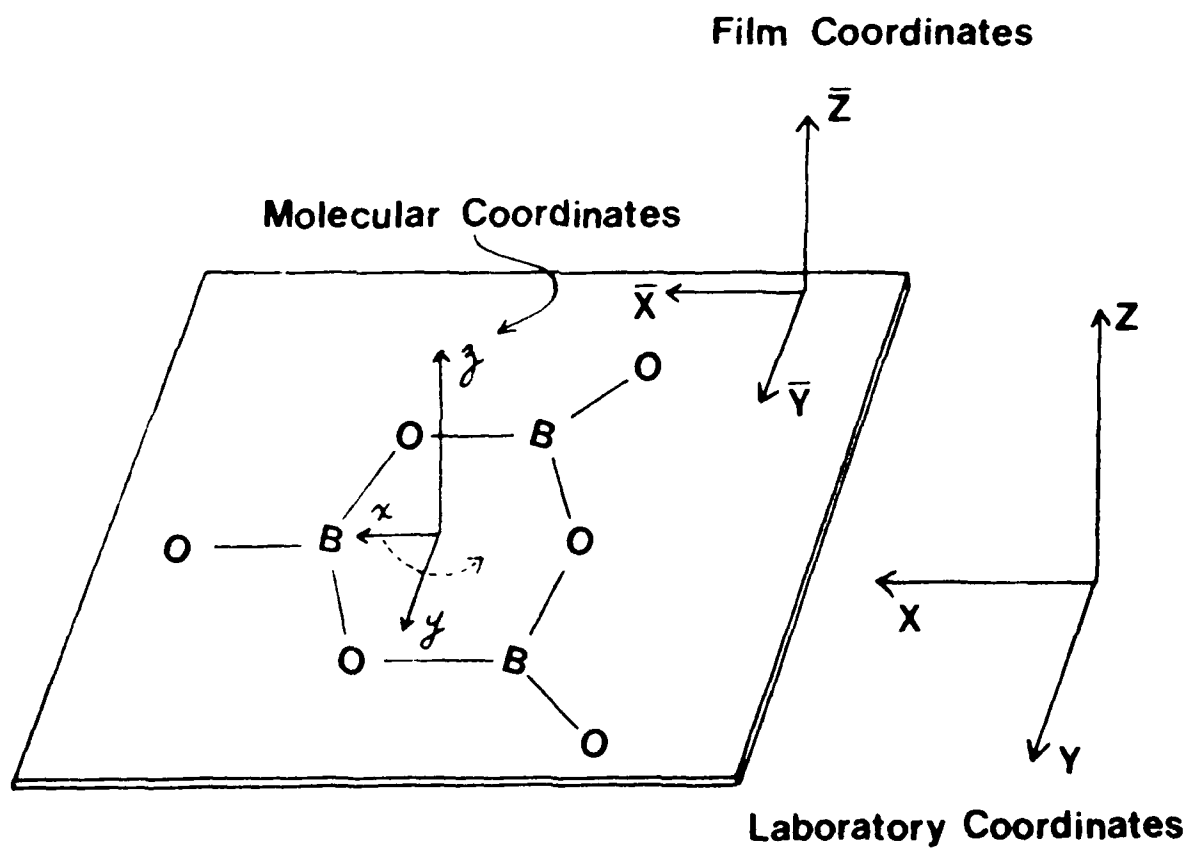
- Figure 1. Representation of the boroxol (B_2O_3) ring structure. The three extra-annular oxygens are bound to trigonal planar B atoms and are the only atoms bound to the ring. The dotted lines from these oxygens to the rest of the network are drawn to emphasize that the extra-annular dihedral angles and those angles with vertices at those oxygens are not defined by the model.
- Figure 2. Definitions of the three-coordinate systems, as noted on the figure.
- Figure 3. Experimental arrangement for Raman scattering. Two film orientations are shown: (a) Tilt 1, $\theta = 45^\circ$; (b) Tilt 2, $\phi = 90^\circ$. Figure (b) also corresponds to Tilt 3, $\theta = 90^\circ$.
- Figure 4. Definitions and representations of the orientations of B_2O_3 films in the three "Tilts". In Tilt 1 (a) θ varies in the ZX plane; in Tilt 2 (b) ϕ varies in the XY plane; and, in Tilt 3 (c) θ varies in the ZY plane.
- Figure 5. Plots of the calculated intensity in the non-zero orientation relative to that in the HH orientation, i/HH , versus tilt angle for the case $a \gg b$. Plot (a) is for Tilt 1, tilt angle is θ , and i/HH is VV/HH . Plot (b) applies to Tilt 2, tilt angle ϕ , and Tilt 3, tilt angle θ , and represents HV/HH in Tilt 2 and VH in Tilt 3.

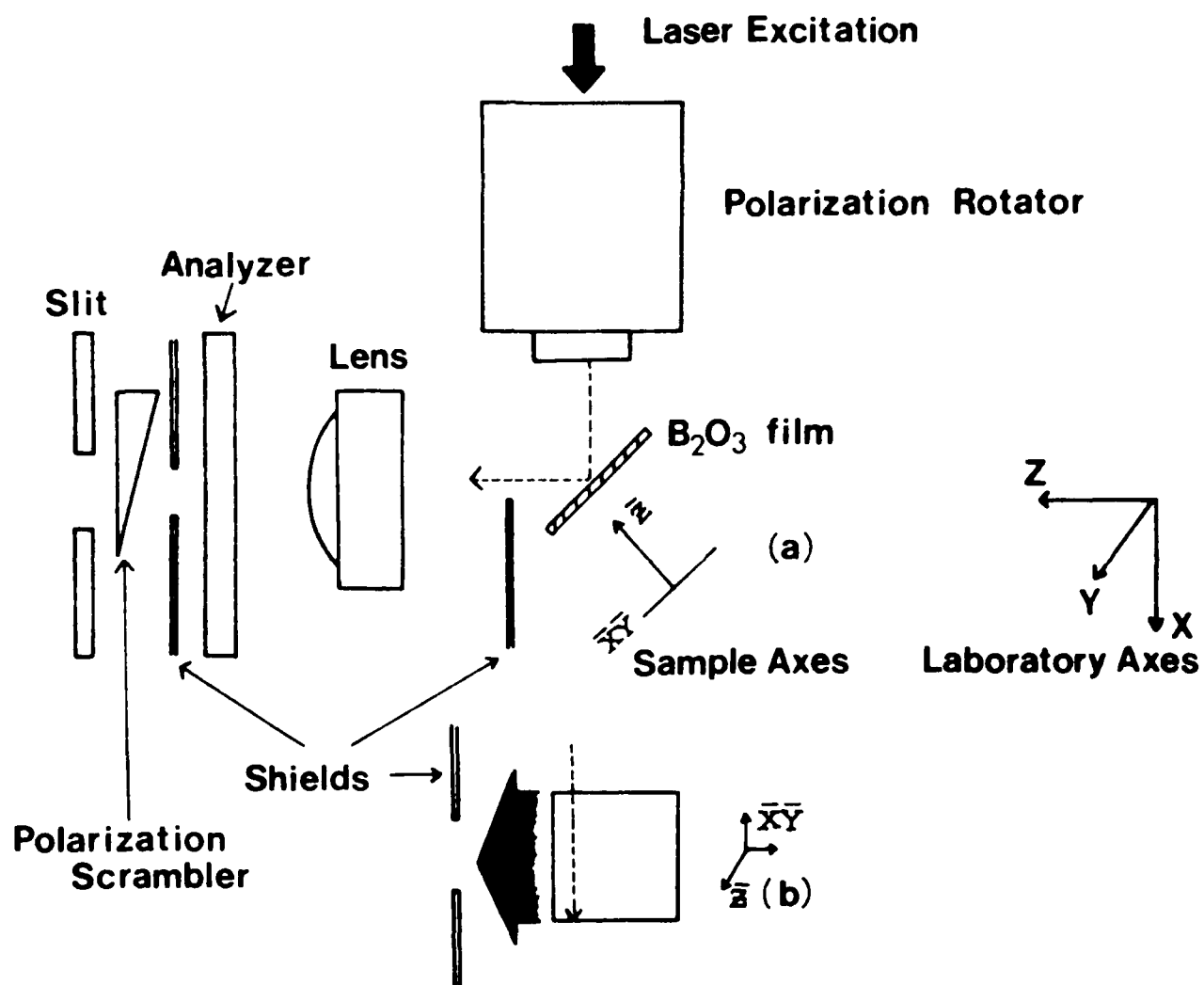
Figure 6. Sample geometry and the directions of the incident and scattered (collected) light in laboratory coordinates. Here, (a) is the bulk (cloven block) sample, and (b) shown the sample mounting used for Tilt 1, $0^\circ \leq \theta < 90^\circ$.

Figure 7. Comparison of the calculated and observed dependence of relative intensity on tilt angle (θ) for Tilt 1. The upper graph refers to the VV/HH ratio, while the lower one applies to the HV/HH and VH/HH ratios.

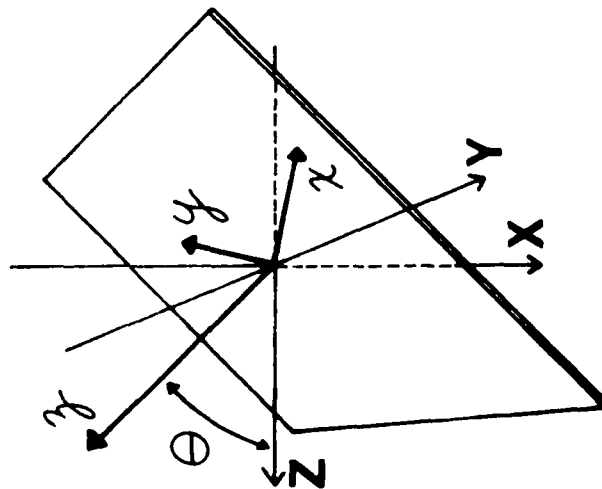


COORDINATE DEFINITIONS

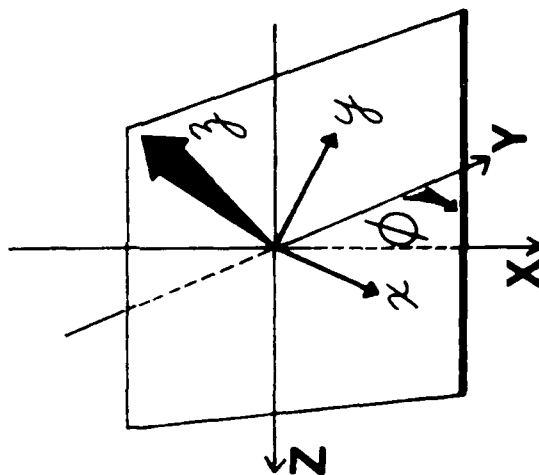




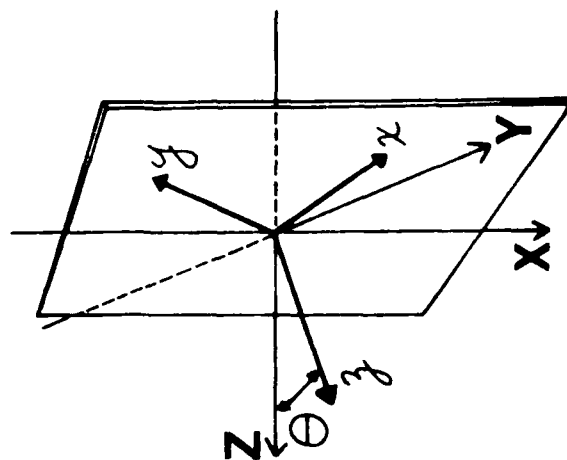
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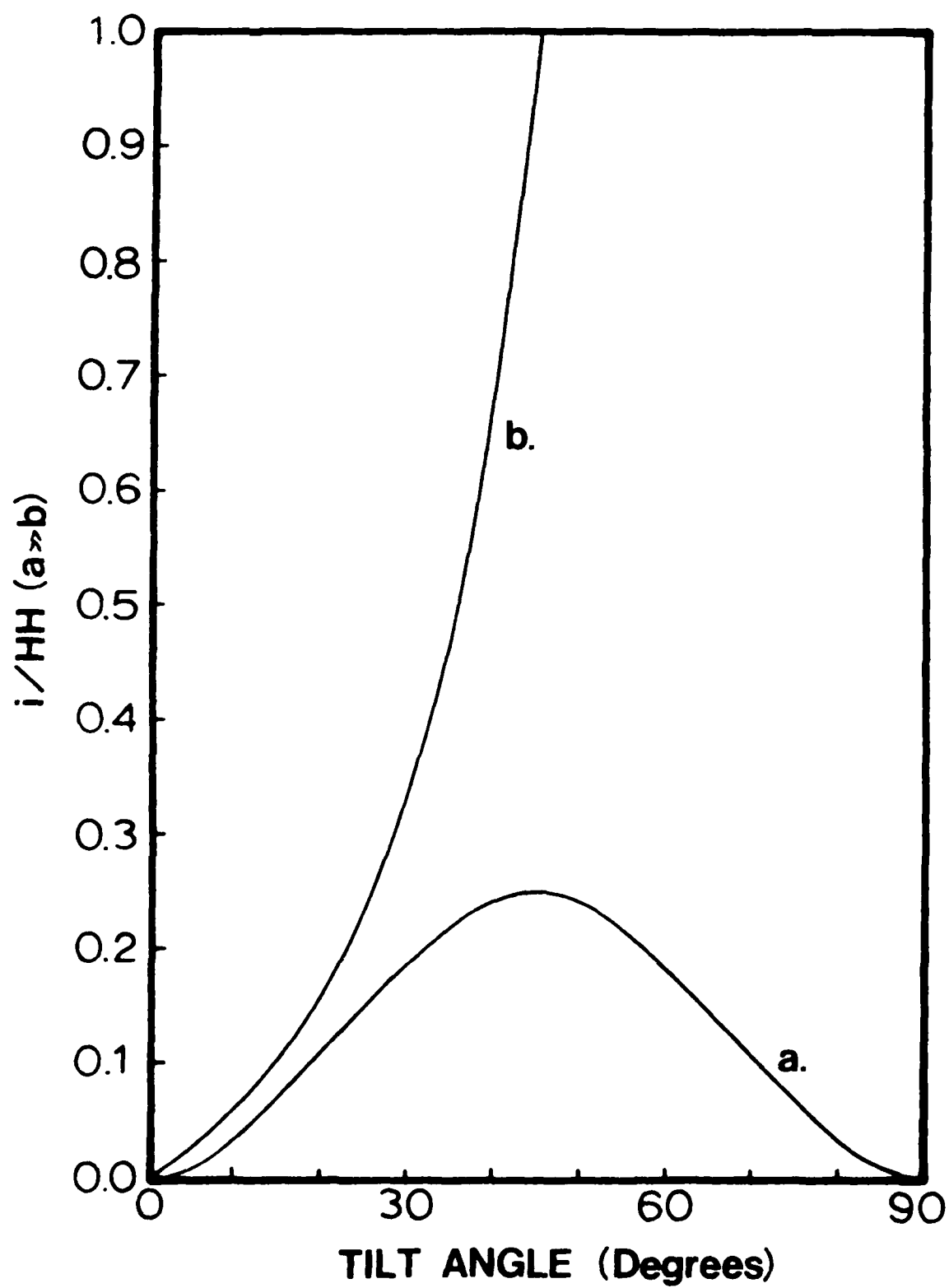


b. TILT 2

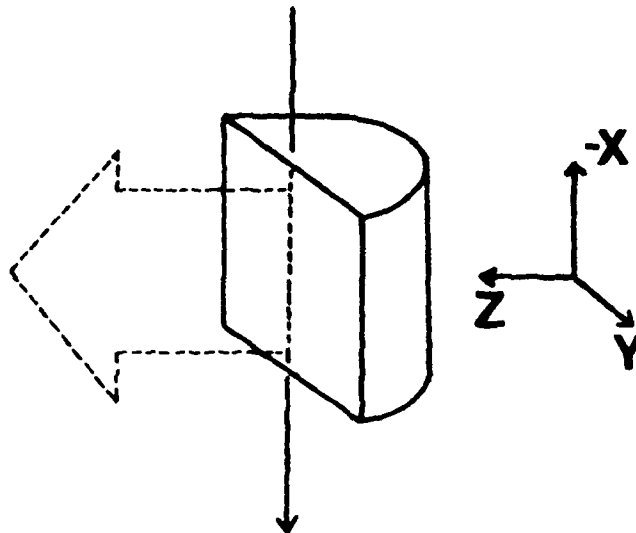


c. TILT 3

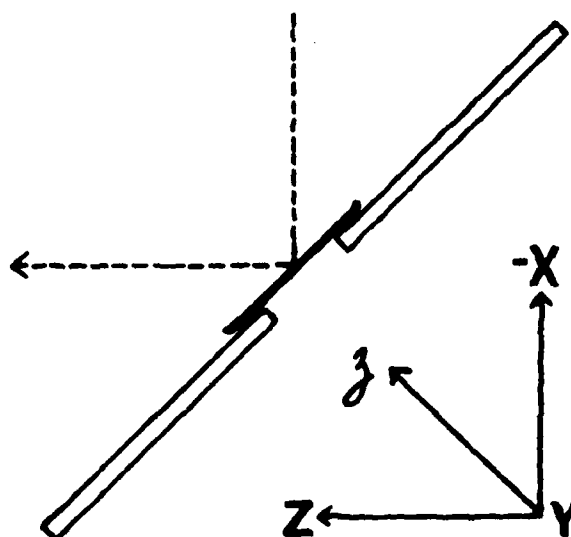


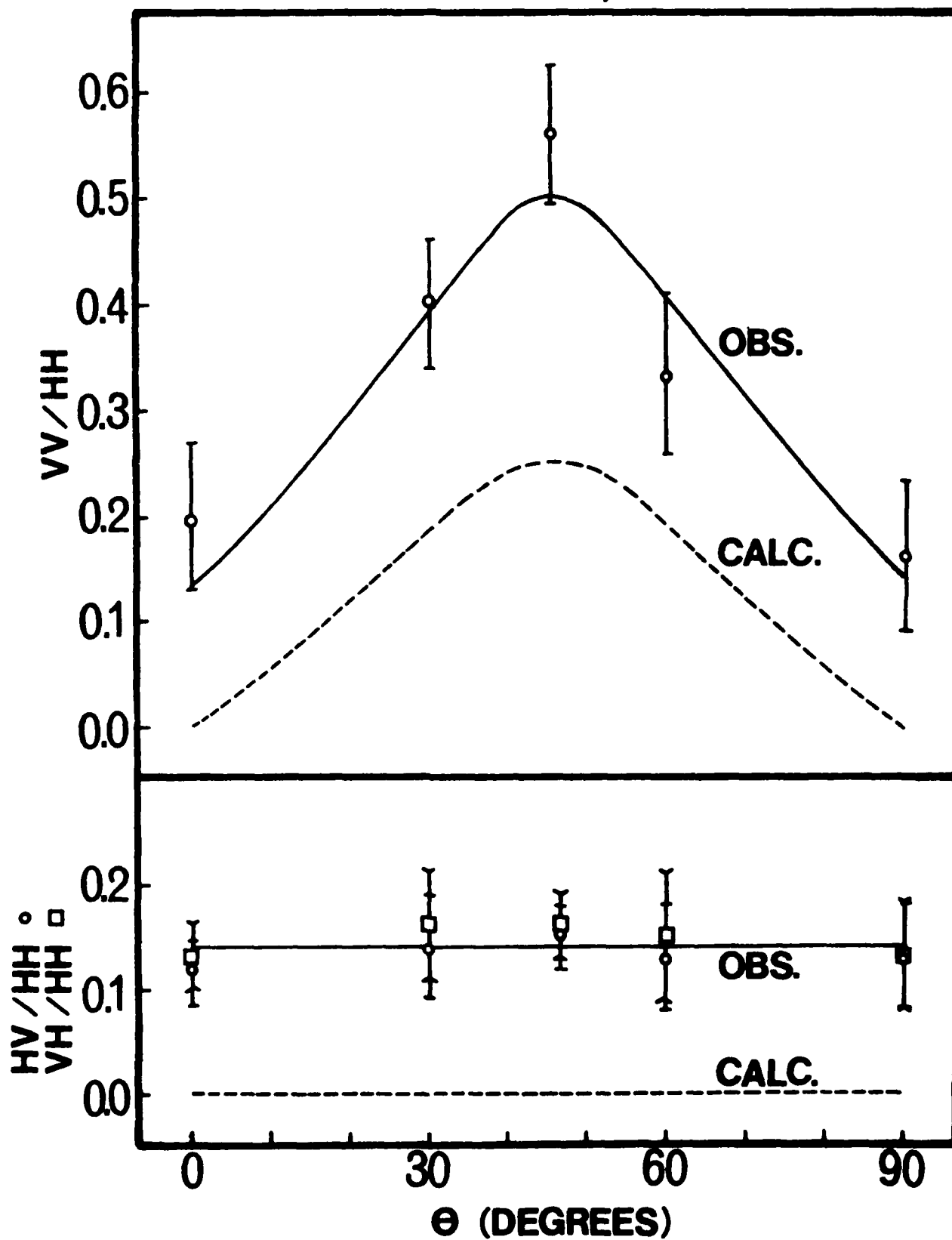


a.



b.



TILT 1 ($\phi = 90^\circ$)

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